

# PS02A -Teach About Water and Solutions

Use with BrishLab [PS02A](#)  
Done By: Coach

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1- How many sides to a container do you need for a liquid and why?

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Para 2

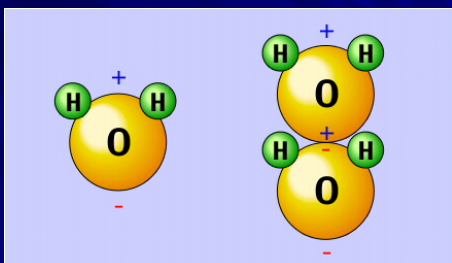


5 out of 6 - no top needed. Gravity holds it in.

[Image Link](#)

2- Explain a polar molecule.

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Para 4



The different ends of a water molecule have charges - like the poles of a magnet.

[Image Link](#)

3- Why is water a "special" molecule?

Use the word "skin".

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Para 5



A paper clip can "float" on the skin of the water.

[Image Link](#)

4- Why does ice float?

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Para 6

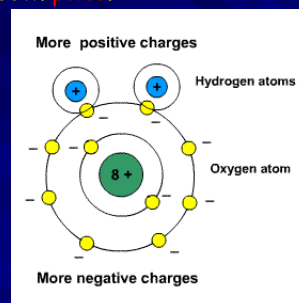


Ice is less dense than water.

[Image Link](#)

5- Why does a water molecule have a slight charge at both poles?

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Para 7

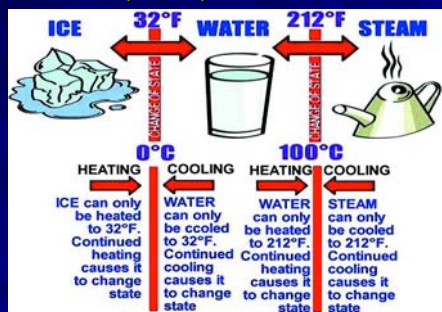


Atoms are not the same on each end of a water molecule so there is a slight charge.

[Image Link](#)

6- What two things happen to water at 100° C (212° F)?

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Para 10

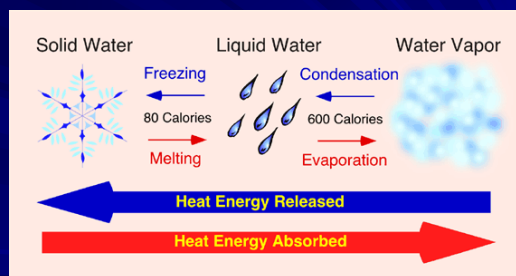


Water either evaporates or condenses at 100° C

[Image Link](#)

7- What happens to water to make it solid and then into a liquid?

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Para 11



Removing heat turns liquid water into solid water (ice). Adding heat turns solid water (ice) into liquid water.

[Image Link](#)

8- Why do we use salt on an icy street?

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Para 12



Salt lowers the freezing point of water so ice does not form.

[Image Link](#)

9- Salt in water is a solution. Salt is the solute and water is the solvent - there is more water than salt.

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Para 12

How can you remember this?



The 'u' in Solute is what 'u' put into the larger solvent.

[Image Link](#)

10- Why is water known as the "Universal Solvent"?

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Para 13

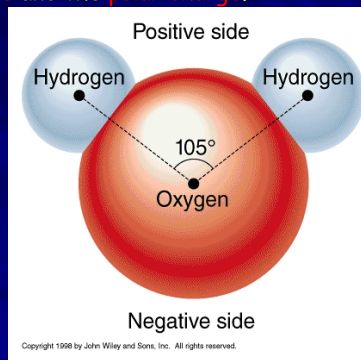


Because so many things can be dissolved into water.

[Image Link](#)

Wrap It Up: - Draw and color a water molecule and the polar charge.

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Para 4



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