PS02A –Teach About

Water and Solutions

Use with BrishLab PS02A
Done By: Coach

1- How many sides to a container do you need for a liquid and why?

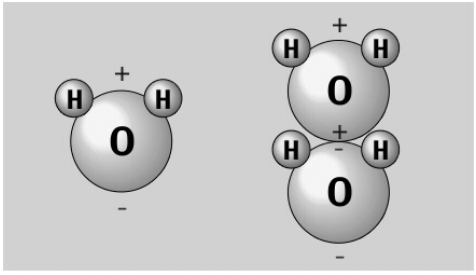
Page 1
Para 2

5 out of 6 – no top needed. Gravity holds it in.

Image Link

2- Explain a polar molecule.

Page 1
Para 4



The different ends of a water molecule have charges – like the poles of a magnet.

Image Link

3- Why is water a "special" molecule? Use the word "skin".

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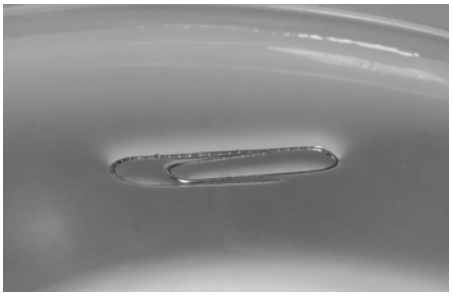
Para 5

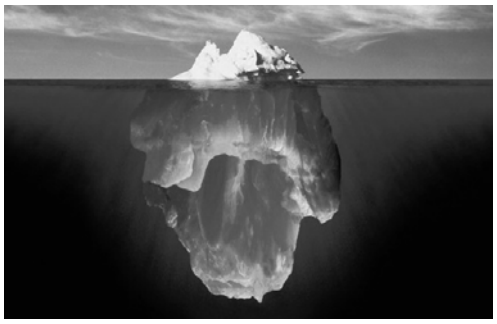


A paper clip can "float" on the skin of the water.

Image Link

4- Why does ice float?

Page 1
Para 6



Ice is less dense than water.

Image Link

5- Why does a water molecule have a slight charge at both poles?

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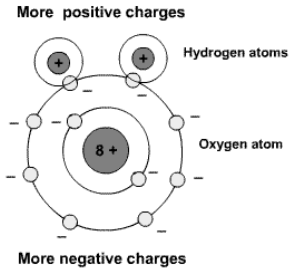
Para 7



Atoms are not the same on each end of a water molecule so there is a slight charge.

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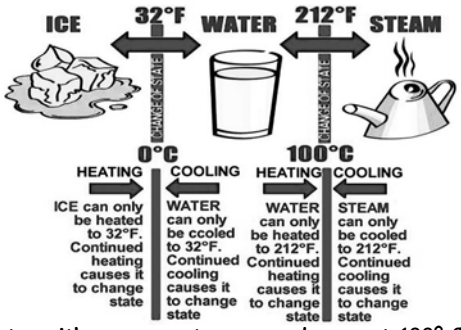
6- What two things happen to water at 100° C (212° F)?

Page 2
Para 10

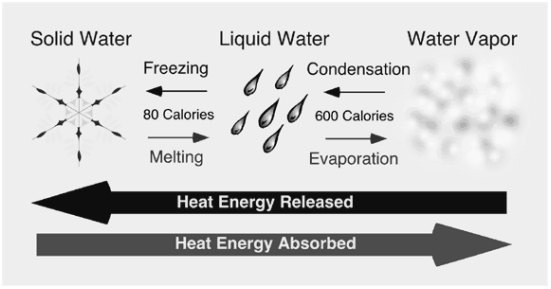


Water either evaporates or condenses at 100° C

Image Link

7- What happens to water to make it solid and then into a liquid?

Page 2
Para 11



Removing heat turns liquid water into solid water (ice). Adding heat turns solid water (ice) into liquid water.

Image Link

8- Why do we use salt on an icy street?

Page 2
Para 12

Salt lowers the freezing point of water so ice does not form.

Image Link

9- Salt in water is a solution. Salt is the solute and water is the solvent – there is more water than salt.

How can you remember this?

Page 2
Para 12

The 'u' in Solute is what 'u' put into the larger solvent.

Image Link

10- Why is water known as the "Universal Solvent"?

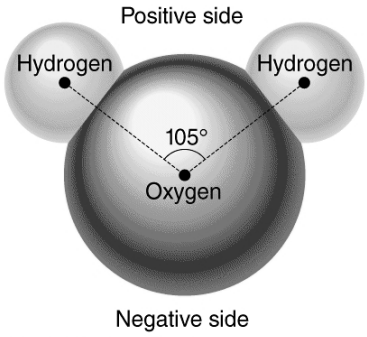
Page 2
Para 13

Because so many things can be dissolved into water.

Image Link

Wrap It Up: - Draw and color a water molecule and the polar charge.

Page 1
Para 4



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